

In spite of these difficulties, extrapolation to 1 atm gives a value of K for sodium fluorenyl at $\sim 25^\circ\text{C}$ which is identical with that obtained by extrapolation to 25°C , of the previously reported data¹ pertaining to lower temperatures (i.e., $K_{25^\circ, 1\text{atm}} \sim 0.058$). Some discrepancy is found, however, for the lithium salt (our $K = 2.7$, reported 4.6).

The equation for K is modified if $\epsilon_{\text{loose}}/\epsilon_{\text{tight}} \neq 1$. Denoting this ratio by γ we find

$$K = (R - r_1)/(\gamma - Rr_2).$$

It is possible that our r_2 is too small, because it was determined from the spectrum of lithium fluorenyl recorded at the highest pressure. This may account for our low value of K for the lithium salt at 1 atm. At higher pressures the error is probably partly compensated by γ being larger than 1.

Finally, we have checked that any reasonable changes of r_1 and r_2 have negligible effect upon the final values of ΔV for the sodium salt, although the results obtained for the lithium salt are less certain.

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² S. Claesson, S. Malmrud and B. Lundgren, *Trans. Faraday Soc.*, preceding paper.

³ G. R. Andersson, *Arkiv Kemi.*, 1963, **20**, 513.

⁴ D. Nicholls, C. Sutphen and M. Szwarc, *J. Phys. Chem.*, 1968, **72**, 1021.

⁵ T. E. Hogen-Esch and J. Smid, *J. Amer. Chem. Soc.*, 1966, **88**, 318.

⁶ T. E. Hogen-Esch and J. Smid, private communication.